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Ying-Ying Zheng

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if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have O_2 16.232 grams mole 21.3 grams 0.66 moles 32 g/mol <u>experiment 4 stoichiometry limiting reagents</u>	web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield yield actual $\times 100$ 260.0 g $\times 100$ 86.01	<u>6.5 limiting reagent and percent yield chemistry libretexts</u> - Aug 15 2023 web learning objectives define and determine theoretical yields actual yields and percent yields identify a limiting

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and 50 0 grams of
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0 grams of iodine v
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with 28 0 grams of
carbon monoxide co
determine the mass
of iodine $i2$ which
could be produced
80 g $i2o5$ 1 mol
 $i2o5$ 1 mol $i2$ xs 1
333 8 g $i2o5$ 1 mol
 $i2o5$ 28 g co 1 mol
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