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nuclear reaction we use an equation that identifies the nuclides involved in the reaction their mass numbers and atomic numbers and the other particles involved in the reaction types of particles in nuclear reactions

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web a nucleus changes into a new element by emitting nuclear

radiations these changes are described using nuclear equations
 example ${}_{86}^{219}\text{Rn} \rightarrow {}_{84}^{215}\text{Po}$

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for atomic nuclei the top number a represents the nucleon number or the mass number nucleon number a total number of protons and neutrons in the nucleus the lower number z represents the proton or atomic number

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 $\text{Li} + \text{N}_2 \rightarrow \text{Li}_3\text{N}$
 $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
 $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $\text{NH}_3 + \text{CuO} \rightarrow \text{Cu} + \text{N}_2 + \text{H}_2\text{O}$
 balance the following equations
 $\text{Fe} + \text{S}$

$\text{Cl}_2 + \text{FeCl}_3 + \text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $\text{As} + \text{NaOH} + \text{Na}_3\text{AsO}_3 + \text{H}_2$

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web sep 12 2022 the half life of the nuclei and thus the material is $t_{1/2} = 693 \times 10^6$ million years 35 a the decay constant is $\lambda = 1.99 \times 10^{-5} \text{ s}^{-1}$ $\lambda = 1.99 \times 10^{-5} \text{ s}^{-1}$ b since strontium 91 has an atomic mass of 90.90 g radiation radioactive decay and nuclear equations - May 24 2022

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answer

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