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water is approximately 1 cal/g°C and the change in temperature is 50°C - 20°C = 30°C plugging these values into the equation we get $q = 75.0 \text{ g} \times 1 \text{ cal/g}^\circ\text{C} \times 30^\circ\text{C}$

$= 2250 \text{ cal}$ therefore 2250 cal is the correct

answer rate this question 2/2

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 enthalpy of reaction is $h \text{ prod } h \text{ react}$ the reactant
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 system and the system moves to a lower energy

potential 3

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