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 web  $m \times c \times \Delta t$  equation 1 where m is the total mass of the solution solute plus  
 solvent c is the specific heat of the solution and  $\Delta t$  is the observed temperature  
 change the specific heat of the solution is generally assumed to be the same as  
 that of  
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 web the heat of reaction  $\Delta H_{\text{soln}}$  is written after the products in units of kJ mol  
 rxn for example sodium hydroxide dissolves exothermically releasing 44.2  
 kilojoules per mole dissolved this process is represented as  $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{aq} + \text{OH}^- \text{aq}$   
 $\Delta H_{\text{soln}} = 44.2 \text{ kJ mol}^{-1}$  rxn write equations to similarly represent the dissolving  
 process for  
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 web hand warmers are usually placed inside the gloves to keep your hands warm the  
 heat keeping your hands warm is formed as a result of an exothermic reaction which  
 occurs in a small packet hand warmers have 5 main ingredients in them salt carbon  
 polypropylene vermiculite water and iron  
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 produces a large temperature change in the form of heat in this laboratory  
 experiment the heat of solution will be calculated from the data for six salts in  
 order to  
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 web background the hand warmers are filled with sodium acetate click the metal  
 disc and you will notice the liquid changes state and becomes a solid in the  
 process heat is released the sodium acetate is a supersaturated solution  
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 web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags  
 seal the coin inside a small bag place this bag inside a larger bag containing  
 supersaturated sodium acetate get this by dissolving as much sodium acetate as  
 possible in hot water this chemical hand warmer is reusable  
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 web the goal of my experiment and research is to evaluate the various number of  
 compounds to be used as hand warmers and locate the ones that can raise the  
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 ruptures the inner pouch releasing the ionic salt into the water of the outer  
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