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october 18 1997 was an american professional
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if 80 0 grams of iodine v oxide i_2o_5 reacts
with 28 0 grams of carbon monoxide co
determine the mass of iodine i_2 which could be
produced b if in the above situation only 0 160
moles of iodine i_2 was produced i what mass of
iodine was produced

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yields key $4 \text{ko}_2 + 2 \text{h}_2\text{o} \rightarrow 4 \text{koh} + 3 \text{o}_2$ g if a
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table given the following reaction $\text{qquad text cu} + 2$

text agno 3 rightarrow 2 text ag text cu no 3 text 2 cu 2agno3 2ag cu no3 2

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web a write a balanced equation for the reaction al o₂ al₂o₃ 4 al 3 o₂ 2 al₂o₃ b determine the theoretical yield of al₂o₃ 160 0 g al x 1 mol al x 2 mol al₂o₃ x 101 964 g al₂o₃ 302 3 g al₂o₃ 26 982 g al 4 mol al 1 mol al₂o₃ c determine the

percent yield yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 na_3po_4 nh_4 3po_4 nano_3 which reactant is
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 co_2 g i_2 g a 80 0 grams of iodine v oxide i_2o_5
reacts with 28 0 grams of carbon monoxide co
determine the mass of iodine i_2 which could be
produced 80 g i_2o_5 1 mol i_2o_5 1 mol i_2 xs 1 333
8 g i_2o_5 1 mol i_2o_5 28 g co 1 mol co

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