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4 6g 0.2 2 16 32 g mol n m mr n  
4 6 32 n 0 144 mol y n l y 0 144  
6 02 1023 y 8 669 molecules

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web the mole is how we relate the unbelievably small atoms and molecules that make something up to the measurable properties such as mass which we may observe in a laboratory setting a proper understanding of the mole concept is essential in order to do any calculations on experimental data

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web aug 15 2019 the mole concept is one of the topics with which you leave your preparation of physical chemistry in class 11 one of the reasons for it being taught earlier is that the concept of mole will be required in almost every other topic of physical chemistry that you study later irrespective of the complexity of that topic

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web we interpret this as two molecules of hydrogen react with one molecule of oxygen to make two molecules of water the chemical equation is balanced as long as the coefficients are in the ratio 2 1 2 for instance this chemical equation is also balanced  
 $100\text{H}_2 + 50\text{O}_2 \rightarrow 100\text{H}_2\text{O}$   
 $100\text{H}_2 + 50\text{O}_2 \rightarrow 100\text{H}_2\text{O}$

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web what is the mole concept the mole concept is a convenient method of

expressing the amount of a substance any measurement can be broken down into two parts the numerical magnitude and the units that the magnitude is expressed in for example when the mass of a ball is measured to be 2 kilograms the magnitude is 2 and the unit is

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web jul 29 2021 chemists need a way of simply determining how many molecules they have in a beaker the mole concept which we will introduce here bridges that gap by relating the mass of

a single atom or molecule in amu to the mass of a collection of a large number of such molecules in grams

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web this experimental approach required the introduction of a new unit for amount of substances the mole which remains indispensable in modern chemical science the mole is an amount unit similar to familiar units like pair dozen gross etc it provides a specific measure of the number of atoms or molecules in a bulk sample of matter

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3 or 4 to 3 significant figures iii  
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figures b 35.61 g of tin reacts  
with exactly 42.60 g of chlorine  
cl<sub>2</sub> to form 78.21 g of tin iv  
chloride sncl<sub>4</sub> i calculate the  
number of moles of tin chloride  
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defined as that amount of  
substance which contains  
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the substance is atomic or  
avogadro's number of  
molecules if the substance is  
molecular 1 mole of carbon  
atoms 6.022 10<sup>23</sup> atoms of  
carbon 1 mole of sodium atom  
6.022 10<sup>23</sup> atoms of sodium  
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concept that allows us to  
connect the mass of a  
substance to the number of  
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ions it contains it is based on  
avogadro's number which  
states that one mole of any  
substance contains 6.022 x 10<sup>23</sup>  
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