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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid-base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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$\text{K}^+ + \text{PO}_4^{3-} + \text{Al}^{3+} + 3\text{OH}^- \rightarrow \text{AlPO}_4 + 3\text{H}_2\text{O}$

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product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq $2\text{nh}_4\text{ aq}$ $\text{cr}_2\text{o}_7^{2-}\text{ aq}$ $\text{ag}_2\text{cr}_2\text{o}_7\text{ s}$ $2\text{nh}_4\text{ aq}$ 2f aq because two $\text{nh}_4\text{ aq}$ equations and formulae edexcel balanced ionic equations - Feb 01 2022

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