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hypoiodite chlorite phosphite percarbonate bromate hyposulfite
permanganate carbonite co_3^{2-} mno_3^- so_5^{2-} bro_2^- co_4^{2-} po_5^{3-} clo_1^- io_2^-
 po_4^{3-} no_2^- so_4^{2-}

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ammonium ion nh_4^+ it has five atoms one nitrogen and four hydrogens that share a charge of 1 the polyatomic ions remain intact and parentheses may be required when using subscripts for example ammonium chloride is nh_4cl and ammonium sulfide is nh_4_2s

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cannot be modified in any way table 3 4 1 3 4 1 lists the ion names and ion formulas of the most common polyatomic ions for example no_3^- no_3^- is the nitrate ion it has one nitrogen atom and three oxygen atoms and an overall 1 charge

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polyatomic ions are ions which consist of more than one atom for example nitrate ion no_3^- contains one nitrogen atom and three oxygen atoms the atoms in a polyatomic ion are usually covalently bonded to one another and therefore stay together as a single charged unit

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compounds with polyatomic ions are named the same way as other ionic compounds just put the name of the cation and the name of the anion together if there is more than one polyatomic ion don't add a number to the name

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