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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3 655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1 1 we could experimentally determine the number of moles of the acid in the solution to be 000731 moles making the experimental molar mass acid base titration chemistry 1210 lab report containing an - Jul 15 2023 web chem 1210 spring 2019 experiment 10 11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid astrong acid with sodium hydroxide a strong base and acetic acid a

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considering the 1 1 stoichiometry of this acid base reaction naoh ag c6h4 cooh cook ag c6h4 coona cook ag h2o l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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