

# Yakushi

Donald Fredrick McCallum



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web the balanced chemical equation for this reaction is  $\text{mg s} + 2\text{hno}_3 \text{ aq} \rightarrow \text{mg no}_3 2 \text{ aq} + \text{h}_2 \text{ g}$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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web mar 4 2018  $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg no}_3 2 + \text{h}_2$  on the left side there are mg 1

magnesium in elemental form  $\text{h}_2$  there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6  
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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate  $\text{mgco}_3$  and two moles of nitric acid  $\text{hno}_3$  react to form one mole of magnesium nitrate  $\text{mg no}_3 2$  one mole of water  $\text{h}_2\text{o}$  and one mole of carbon dioxide  $\text{co}_2$  show chemical structure image reactants

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balancing nitrogen by multiplying  $\text{hno}_3$  by 2 we get  $\text{mg}$  s magnesium 2  $\text{hno}_3$  l nitric acid  $\text{mg}$  no 3 2 s magnesium nitrate  $\text{h}_2$  g hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen

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balanced total equation  $\text{mg}$   $\text{oh}$  2 aq 2  $\text{hno}_3$  aq 2  $\text{h}_2$  o  $\text{mg}$  no 3 2 aq total ionic equation  $\text{mg}$  2 2  $\text{oh}$  1 2  $\text{h}$  1 2  $\text{no}_3$  1 2  $\text{h}_2$  o  $\text{mg}$  2 2  $\text{no}_3$  1 net ionic equation 2  $\text{oh}$  1 2  $\text{h}$  1 2  $\text{h}_2$  o sulfuric acid and lithium hydroxide balanced total equation  $\text{h}_2$  so 4 aq 2  $\text{lioh}$  aq 2  $\text{h}_2$

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