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form one mole of aqueous
magnesium nitrate $\text{mg}(\text{no}_3)_2$ and
one mole of dihydrogen h_2 gas
show chemical

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equations our goal is to have the
same number of each type of
atom on both sides of the equation
only change the coefficients these
are the numbers in front
substances never change the
subscripts the small numbers

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web mar 4 2018 $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg}(\text{no}_3)_2 + \text{h}_2$ on the left side the are
 mg 1 magnizium in elemental
form h_2 there are 2 molecules of
nitric acid each acid has one
hydrogen so two molecules of
nitric acid is 2 hydrogens n_2
there are 2 molecules of nitric
acid each acid has one nitrogen so
two molecules of nitric acid is 2
nitrogens o_6

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magnesium hydroxide express
your answer as a chemical
equation including phases 2 enter
a net ionic equation for the
reaction between nitric acid and
magnesium hydroxide express
your answer as a chemical
equation including

mgco₃ hno₃ mg no₃ 2 h₂o co₂

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nitric acid magnesium

magnesium nitrate hydrogen gas

type of chemical reaction for this

reaction we have a single

displacement reaction balancing

strategies in this single

displacement reaction the mg

replaces the h in the hno3

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we ll balance the equation mg

hno3 mg no3 2 no2 h2o and

provide the correct coefficients

for each compound to balance mg

hno3 mg no3 2 no2 h2o you ll

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acid balanced total equation mg

oh 2 aq 2 hno 3 aq 2 h 2 o mg no

3 2 aq total ionic equation mg 2 2

oh 1 2 h 1 2no 3 1 2h 2 o mg 2 2

no 3 1 net ionic equation 2 oh 1 2

h 1 2h 2 o sulfuric acid and

lithium hydroxide balanced total

equation h 2 so 4 aq 2 lihoh aq 2 h

2

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